

- If ω is an imaginary cube root of unity, then $(1+\omega-\omega^2)^7$ equals
 (A) 128ω (B) -128ω (C) $128 \omega^2$ (D) $-128 \omega^2$
- If $A = \begin{pmatrix} \cos \theta & -\sin \theta \\ \sin \theta & \cos \theta \end{pmatrix}$, then which of the following statement is not correct?
 (A) A is orthogonal matrix (B) A' is orthogonal matrix
 (C) Determinant $A=1$ (D) A is not invertible
- A particle of mass 3 kg moves under a force of $(4i+8j+10k)$ Newton. If the particle starts from rest and was at origin initially, what are it new coordinates after 3 seconds?
 (A) (2,4,5) (B) (4,8,10) (C) (6,12,15) (D) (10,0,0)
- A body completes one round of a circle of radius R in 20 seconds. The ratio of displacement to distance after 10 seconds is
 (A) 7 : 11 (B) 10 : 1 (C) 11 : 7 (D) 1 : 10
- A body falls freely under gravity. The distance covered in the first, second and third second of motion are in the ratio
 (A) 1 : 2 : 3 (B) 1 : 4 : 9 (C) 1 : 3 : 5 (D) 2 : 4 : 6
- Which of the waves does **not** belong to electromagnetic wave spectrum?
 (A) X-rays (B) Visible light
 (C) Sound waves (D) Infra-red rays
- Which is the essential condition for producing stationary Interference pattern due to two sources of light?
 (A) Coherent (B) Incoherent
 (C) Monochromatic (D) Both (A) and (C)
- Bending of light around a obstacle is known as
 (A) diffraction (B) reflection
 (C) polarization (D) none of the above
- Mirage formation is due to the variation of layers of
 (A) refractive index (B) colour
 (C) frequency (D) none of the above

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10. A hollow metal sphere of radius 5 cm is charged such that the potential on its surface is 10 volts. The potential at the centre of the sphere is
- (A) zero
 (B) 10 V
 (C) same as at a point 5 cm away from the surface
 (D) same as at a point 25 cm away from the surface
11. A charge q is placed at the centre of the line joining two equal charges Q . The system of three charges will be in equilibrium if q is equal to
- (A) $-Q/4$ (B) $-Q/2$ (C) $+Q/4$ (D) $+Q/2$
12. An electron of mass m_e , initially at rest, moves through a certain distance in a uniform electric field in time t_1 . A proton of mass m_p , also, initially at rest, takes time t_2 to move through an equal distance in this uniform electric field. Neglecting the effect of gravity, the ratio t_2/t_1 is nearly equal to
- (A) 1.0 (B) $(m_p/m_e)^{0.5}$ (C) $(m_e/m_p)^{0.5}$ (D) 1836
13. Sodium has 11 electrons. If the sequence in which the energy levels are filled is 1s, 2s, 2p, 3s, 3p, 4s, 3d,.... the ground state of sodium is
- (A) $^3P_{1/2}$ (B) $^2P_{1/2}$ (C) $^1P_{1/2}$ (D) $^3S_{1/2}$
14. The maximum kinetic energy of electrons in a photoelectric effect depends on
- (A) intensity of incident light (B) frequency of the incident light
 (C) polarization of the incident light (D) angle of incidence
15. Electromagnetic radiations will be emitted in the case of a
- (A) neutron moving in a straight line with a constant speed
 (B) proton is moving in straight line with a constant speed
 (C) proton moving in a circle with a constant speed
 (D) electron moving in a straight line with a constant speed
16. The de Broglie wave length of an electron of Kinetic energy 500 eV is
- (A) 14.82 Å (B) 24.82 Å (C) 34.82 Å (D) 44.82 Å
17. The maximum number of electrons in an atomic orbit of principal quantum number $n = 3$ is
- (A) 3 (B) 6 (C) 9 (D) 18

18. A sensitive galvanometer is converted into an ammeter by connecting a _____ resistance in _____ with its coil
- (A) Low, parallel (B) Low, series
(C) High, parallel (D) High, series
19. The intensity of the magnetic field H at a point on the axial line of a bar magnet is
- (A) $2Md/(d^2-l^2)^{1/2}$ (B) $2Md/(d^2-l^2)^2$ (C) $2Md/(d^2-l^2)$ (D) $2Md/(d^2-l^2)^3$
20. If a particle is moving in a uniform magnetic field, then
- (A) its momentum changes but total energy remains same
(B) both momentum and total energy remains the same
(C) its total energy changes but momentum remains same
(D) both momentum and total energy will change
21. $M L^2 T^{-2}$ is the dimensional formula of
- (A) angular momentum (B) linear momentum
(C) angular velocity (D) none of the above
22. Which of the following quantity has the same dimensions as the latent heat?
- (A) Work per unit mass (B) Specific heat per unit mass
(C) Force per unit velocity (D) Acceleration per unit displacement
23. An object is projected upwards with a velocity of 10 m/sec. If $g = 10 \text{ m/sec}^2$, then it will strike the ground in approximately
- (A) 1 sec (B) 2 sec (C) 3 sec (D) 5 sec
24. A passenger (facing the direction of motion) tosses a coin in a train. If the coin falls behind him, the train must be moving with
- (A) an acceleration (B) retardation
(C) uniform speed (D) any of these
25. An electron in Bohr's orbit has a constant (i) acceleration (ii) Momentum (iii) force
- (A) (i) only (B) (ii) only (C) (iii) only (D) none of these
26. Kinetic energy per gram molecule is given by
- (A) $3/2 RT$ (B) $3/2 KT$ (C) $5/2 KT$ (D) $1/2 KT$

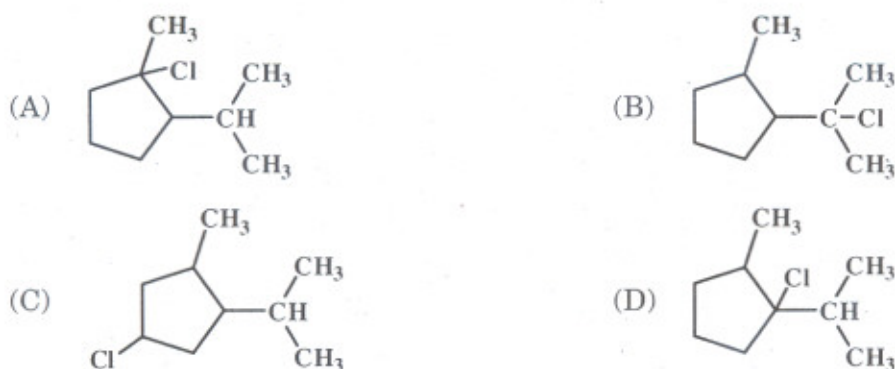
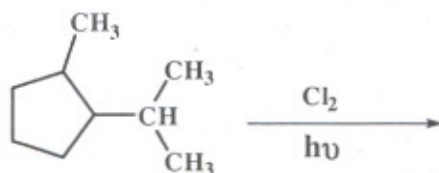
27. A gas will behave as an ideal gas at
(A) High pressure and low temperature
(B) High temperature and high pressure
(C) At very low pressure and high temperature
(D) None of these
28. The resistance of a device dropped drastically when the polarity of the meter changed. The device could be a
(A) Resistor (B) Inductor (C) Capacitor (D) Diode
29. The half life period of neutron is 13 minutes approximately. The intensity of neutron beam traveling in free space with a velocity of 30 km per sec is reduced to half when it cover a distance of
(A) 390000 km (B) 11700 km (C) 46800 km (D) 23400 km
30. Magnetic field does not cause deflection in
(A) γ -rays (B) β^- -rays (C) β^+ -rays (D) α -rays
31. Which of the following represents the possible oxidation states of iodine in its compounds?
(A) -1, 0 (B) -1, +1, +2, +3, +4, +5, +6, +7
(C) -1, +1, +3, +5, +7 (D) -1, -3, -5, -7
32. Which of the following metal ion is essential for blood coagulation?
(A) Iron (B) Phosphorous
(C) Potassium (D) Calcium
33. Oxide of which of the elements among Mn, Pb, S, and Ba in its highest oxidation state will be acidic as well as oxidizing?
(A) Sulphur (B) Manganese (C) Lead (D) Barium
34. The hybridization and geometry of the molecule XeOF_2 is
(A) sp^3d , T shaped (B) sp^3 , tetrahedral
(C) dsp^2 , square planar (D) sp^2 , trigonal planar

35. A nuclear reaction between ^{14}N and ^4He produces ^{17}O as one of the products. The balanced equation for the reaction is
- (A) $^{14}\text{N} + ^4\text{He} \longrightarrow ^{17}\text{O} + ^1\text{p} + \gamma$
 (B) $^{14}\text{N} + ^4\text{He} \longrightarrow ^{17}\text{O} + ^1\text{n} + \beta$
 (C) $^{14}\text{N} + ^4\text{He} \longrightarrow ^{17}\text{O} + ^1\text{n} + \gamma$
 (D) $^{14}\text{N} + ^4\text{He} \longrightarrow ^{17}\text{O} + \alpha + \gamma$
36. Which among the following halides is not a Lewis acid?
 (i) CCl_4 (ii) BCl_3 (iii) SnCl_2 (iv) InBr_3
 (A) (ii) (B) (i) (C) (iv) (D) (iii)
37. Third ionization energy of titanium is required to carry out which of the following processes?
 (A) $\text{Ti}^{3+}(\text{g}) + \text{e}^- \rightarrow \text{Ti}^{2+}(\text{g})$ (B) $\text{Ti}^{2+}(\text{g}) \rightarrow \text{Ti}^{3+}(\text{g}) + \text{e}^-$
 (C) $3\text{Ti}(\text{g}) \rightarrow \text{Ti}^+(\text{g}) + 3\text{e}^-$ (D) $\text{Ti}(\text{g}) \rightarrow \text{Ti}^{3+}(\text{g}) + 3\text{e}^-$
38. Which of the following ion has longest radius?
 (A) Ca^{2+} (B) K^+ (C) S^{2-} (D) Cl^-
39. An element present in 6th period 3rd group has an atomic number 57. The atomic number of the element present in 6th period 4th group is
 (A) 58 (B) 72 (C) 71 (D) 56
40. Which of the following statement is correct with respect to the formula and name of binary compound between oxygen and fluorine?
 (A) F_2O , fluorine oxide (B) OF_2 , Oxygen fluoride
 (C) OF , Oxygen monofluoride (D) FO , Fluorine monoxide
41. Which of the following halide will readily undergo hydrolysis?
 (A) CCl_4 (B) SiCl_4 (C) NaCl (D) HCl
42. Which of the following compounds would have the highest boiling point?
 (A) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$ (B) CH_3NH_2
 (C) CH_3OH (D) CH_2F_2

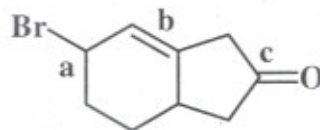
43. What is the total number of sigma bonds found in the following compound?



- (A) 8 (B) 10 (C) 11 (D) 15
44. Which one of the following compounds represents the major monochlorination isomer formed in the following reaction?

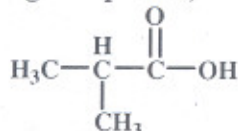


45. The major reason that phenol is a better Brønsted acid than cyclohexanol is
- (A) it is a better proton donor.
- (B) the cyclohexyl group is an electron donating group by induction, which destabilizes the anion formed in the reaction
- (C) phenol is able to stabilize the anion formed in the reaction by resonance
- (D) the phenyl group is an electron withdrawing group by induction, which stabilizes the anion formed in the reaction.
46. Which of the functional groups on the following molecule are susceptible to nucleophilic attack?



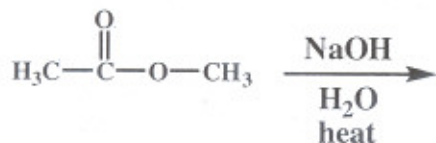
- (A) "a" and "b" (B) "a" and "c"
- (C) "b" and "c" (D) "b" only

47. The best synthesis of the following compound, 2-methylpropanoic acid, will be...



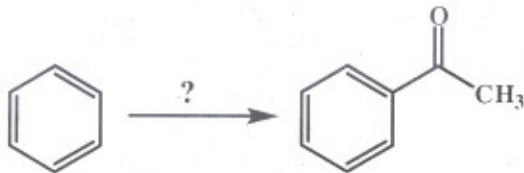
- (A) $\begin{array}{c} \text{H} \\ | \\ \text{H}_3\text{C}-\text{C}-\text{Br} \\ | \\ \text{CH}_3 \end{array} \xrightarrow{\text{KCN}} \xrightarrow[\text{H}_2\text{O, heat}]{\text{NaOH}}$
- (B) $\begin{array}{c} \text{H}_3\text{C}-\text{C}=\text{CH}_2 \\ | \\ \text{CH}_3 \end{array} \xrightarrow[\text{H}^+, \text{cold}]{\text{KMnO}_4}$
- (C) $\begin{array}{c} \text{H} \\ | \\ \text{H}_3\text{C}-\text{C}-\text{Br} \\ | \\ \text{CH}_3 \end{array} \xrightarrow[\text{Et}_2\text{O}]{\text{Mg}} \xrightarrow[\text{H}_2\text{O, heat}]{\text{H}_2\text{C}=\text{O}} \xrightarrow[\text{H}^+, \text{cold}]{\text{H}_3\text{O}^+} \xrightarrow{\text{KMnO}_4}$
- (D) $\begin{array}{c} \text{H} \\ | \\ \text{H}_3\text{C}-\text{C}-\text{Br} \\ | \\ \text{CH}_3 \end{array} \xrightarrow[\text{Et}_2\text{O}]{\text{Mg}} \xrightarrow{\text{CO}_2} \xrightarrow[\text{H}_3\text{O}^+]{}$

48. Which of the following is an intermediate for the basic hydrolysis of methyl ethanoate?



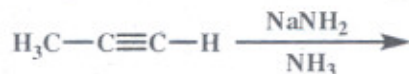
- (A) $\begin{array}{c} \text{O} \\ \parallel \\ \text{H}_3\text{C}-\text{C}-\text{O}^+-\text{CH}_3 \\ | \\ \text{H} \end{array}$
- (B) $\begin{array}{c} \text{OH} \\ | \\ \text{H}_3\text{C}-\text{C}-\text{O}-\text{CH}_3 \\ | \\ \ominus \end{array}$
- (C) $\begin{array}{c} \text{OH} \\ | \\ \text{H}_3\text{C}-\text{C}-\text{O}-\text{CH}_3 \\ | \\ \text{O}^- \end{array}$
- (D) $\begin{array}{c} \text{OH} \\ | \\ \text{H}_3\text{C}-\text{C}-\text{O}-\text{CH}_3 \\ | \\ \text{OH}_2^+ \end{array}$

49. Which of the following reagents would be the best reactants for the following reaction?



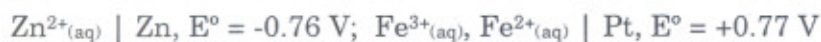
- (A) $\begin{array}{c} \text{H}_2 \\ | \\ \text{H}_3\text{C}-\text{C}-\text{Cl} \\ | \\ \text{AlCl}_3 \end{array}$
- (B) $\begin{array}{c} \text{O} \\ \parallel \\ \text{H}_3\text{C}-\text{C}-\text{OH} \\ | \\ \text{NaOH} \end{array}$
- (C) $\begin{array}{c} \text{O} \\ \parallel \\ \text{H}_3\text{C}-\text{C}-\text{Cl} \\ | \\ \text{AlCl}_3 \end{array}$
- (D) $\begin{array}{c} \text{O} \\ \parallel \\ \text{H}_3\text{C}-\text{C}-\text{H} \\ | \\ \text{AlCl}_3 \end{array}$

50. What is the major product expected from the following reaction?



- (A) $\text{H}_3\text{C}-\text{C}\equiv\text{C}^{\ominus} \text{Na}^{\oplus}$ (B) $\text{H}-\text{C}\equiv\text{C}-\text{CH}_2^{\ominus} \text{Na}^{\oplus}$
- (C) $\text{H}_3\text{C}-\overset{\text{NH}_2}{\text{C}}=\text{CH}_2$ (D) $\text{H}_2\text{N}-\overset{\text{H}_2}{\text{C}}-\text{C}\equiv\text{C}-\text{H}$
51. The half-life of radioactive Palladium-233 is 28 days. How many days will it take for the radioactivity to fall to one eighth of the initial value?
- (A) 84 (B) 28 (C) 112 (D) 56
52. One property of a buffer solution, prepared from a weak acid and its sodium salt, is that
- (A) its pH is less than the pH of the original acid.
 (B) its pH is unaffected by the addition of any quantity of H^+ ions.
 (C) its pH is greater than the pH of the original acid.
 (D) it has a pH of 7
53. The following are standard enthalpies of formation:
 $\text{CO}_{(\text{g})}$: -110 kJ/mole $\text{CO}_{2(\text{g})}$: -394 kJ/mole $\text{H}_2\text{O}_{(\text{g})}$: -242 kJ/mole
 What is the standard enthalpy change for the following reaction?
 $\text{CO}_{(\text{g})} + \text{H}_2\text{O}_{(\text{g})} \rightarrow \text{CO}_{2(\text{g})} + \text{H}_{2(\text{g})}$
- (A) -262 kJ/mole (B) +262 kJ/mole (C) -42 kJ/mole (D) +42 kJ/mole
54. A real gas fails to obey the ideal gas equation under all conditions because
- I. the molecules do not all have the same velocity.
 II. collisions are not perfectly elastic.
 III. the molecules have a finite size.
- (A) I and II are correct (B) II and III are correct
 (C) I is the only correct response (D) III is the only correct response
55. Which of the following is/are polar?
- I. CH_2Cl_2 II. CHCl_3 III. CCl_4
- (A) I and II are correct (B) II and III are correct
 (C) I is the only correct response (D) III is the only correct response

56. From the following data :



it can be deduced that

I. the standard E.M.F. for the cell $\text{Zn} | \text{Zn}^{2+}_{(\text{aq})} ; \text{Fe}^{3+}_{(\text{aq})}, \text{Fe}^{2+}_{(\text{aq})} | \text{Pt}$ is 0.01 V.

II. zinc is a more powerful reductant than Fe^{2+} ions.

III. Fe^{3+} can oxidize zinc under standard conditions.

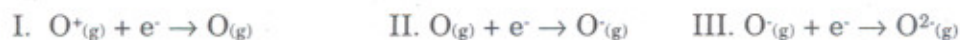
(A) I and II are correct

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(C) I is the only correct response

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57. For which of the following is ΔH° positive?



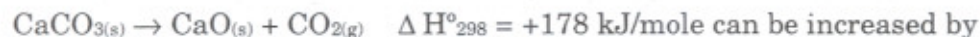
(A) I and II are correct

(B) II and III are correct

(C) I is the only correct response

(D) III is the only correct response

58. The amount of carbon dioxide present at equilibrium in the reaction



I. raising the temperature

II. raising the pressure

III. adding some more solid calcium oxide

(A) I and II are correct

(B) II and III are correct

(C) I is the only correct response

(D) III is the only correct response

59. The Faraday constant is

(A) 96500 coulomb

(B) 95600 coulomb

(C) 96500 ohm

(D) 6.023×10^{23}

60. The Bragg's equation for X-ray diffraction by a crystal is

(A) $n\lambda = 2d \sin\theta$

(B) $n\lambda = 2d \sec\theta$

(C) $n\lambda = 2d \theta$

(D) $n\lambda = 2d \sin\theta$

61. The product of three positive reals is 1 and their sum is greater than sum of their reciprocals. Exactly one of them is greater than

(A) 0

(B) 1

(C) -1

(D) -2

62. If a, b, c are in G.P., then a + b, 2b, b + c are in

(A) A.P.

(B) G.P

(C) H.P

(D) Both A.P and G.P

63. If the fourth roots of unity are z_1, z_2, z_3, z_4 , then $z_1^2 + z_2^2 + z_3^2 + z_4^2$ is equal to
 (A) 1 (B) 0
 (C) i (D) none of these
64. If z and ω are two non zero complex numbers such that $|z\omega| = 1$, and $\arg(z) - \arg(\omega) = \frac{\pi}{2}$, then $\bar{z}\omega$ is equal to
 (A) 1 (B) -1 (C) i (D) $-i$
65. The value of $(1+i)^3 + (1+i)^6$ is
 (A) i (B) $2(-1+5i)$ (C) $1-5i$ (D) $1+5i$
66. If $(x-1)^4 - 16 = 0$, then the sum of non-real complex values of x is
 (A) 2 (B) 0 (C) 4 (D) 6
67. The value of $\log_3 5 \cdot \log_{25} 27$ is
 (A) 3 (B) $3/2$ (C) $1/2$ (D) 1
68. The curve represented by $\operatorname{Re}(z^2) = 4$ is
 (A) a parabola (B) an ellipse
 (C) a circle (D) a rectangular hyperbola
69. In a class tournament where the participants were to play one game with another, two class players fell ill, having played 3 games each. If the total number of games played is 84, the number of participants at the beginning was
 (A) 22 (B) 15
 (C) 17 (D) none of the above
70. 15 identical balls have to be put in 5 different boxes. Each box can contain any number of balls. Total number of ways of putting the ball into the box so that each box contain at least 2 balls, is equal to
 (A) 9C_5 (B) ${}^{10}C_5$ (C) 6C_5 (D) ${}^{10}C_6$
71. For $2 \leq r \leq n$, $\binom{n}{r} + 2\binom{n}{r-1} + \binom{n}{r-2} =$
 (A) $\binom{n+1}{r-1}$ (B) $2\binom{n+1}{r+1}$ (C) $2\binom{n+2}{r}$ (D) $\binom{n+2}{r}$

72. If A, B are square matrices, such that $A^2 = A$, $B^2 = B$ and A, B commute then:
 (A) $(AB)^2 = I$ (B) $(AB)^2 = AB$ (C) $(AB)^2 = 0$ (D) $(AB)^2 = 2$
73. Of three independent events the chance that only the first occurs is a, that only the second occurs is b, and only the third occurs is c. Then the probability of occurrence of these three events are (where x is a root of the equation), $(a+x)(b+x)(c+x) = x^2$, is
 (A) $\frac{a}{a+x}, \frac{b}{b+x}, \frac{c}{c+x}$ (B) $\frac{b}{a+x}, \frac{a}{b+x}, \frac{c}{c+x}$
 (C) $\frac{c}{a+x}, \frac{b}{b+x}, \frac{a}{c+x}$ (D) $\frac{c}{a+x}, \frac{b}{b+x}, \frac{ab}{c+x}$
74. Six faces of a die are marked with the numbers 1, -1, 0, -2, 2 and 3. The die is thrown thrice. The probability that the sum of the numbers thrown is six, is
 (A) $\frac{1}{72}$ (B) $\frac{1}{12}$ (C) $\frac{5}{108}$ (D) $\frac{1}{36}$
75. If ABC is a triangle and $\tan \frac{A}{2}, \tan \frac{B}{2}, \tan \frac{C}{2}$ are in H.P, then the minimum value of $\cot \frac{B}{2}$ is equal to
 (A) $-\sqrt{3}$ (B) $\sqrt{3}$ (C) 2 (D) -2
76. The value of y so that the line through (3,y) and (2,7) is parallel to the line through (-1,4) and (0,6) is
 (A) 5 (B) -5 (C) 9 (D) -9
77. Area of triangle formed by the lines $x + y = 3$ and angle bisector of the pair of straight lines $x^2 + y^2 + 2y = 1$ is
 (A) 2 sq. units (B) 4 sq. units (C) 6 sq. units (D) 8 sq. units
78. Lines $3x+4y+6=0$, $\sqrt{2}x+\sqrt{3}y+2\sqrt{2}=0$ and $4x+7y+8=0$
 (A) Concurrent (B) Parallel
 (C) Sides of a triangle (D) None of these
79. Let $f(x) = 4$ and $f'(x) = 4$, then $\lim_{x \rightarrow 2} \frac{xf(2) - 2f(x)}{x-2}$ equals to
 (A) 2 (B) -2 (C) -4 (D) 3
80. $\lim_{x \rightarrow \infty} \left(\frac{x^2 + 5x + 3}{x^2 + x + 3} \right)^x$ is equal to
 (A) e^4 (B) e^2 (C) e^3 (D) e

81. $\lim_{x \rightarrow 0} \frac{\tan x - \sin x}{x^3}$ is equal to
 (A) 0 (B) 1 (C) 1/2 (D) -1/2
82. Let $f : (0, +\infty) \rightarrow R$ and $F(x) = \int_0^x f(t) dt$. If $F(x^2) = x^2(1+x)$, then $f(4)$ equals
 (A) 5/4 (B) 7 (C) 4 (D) 2
83. The function $f(x) = \frac{x}{2} + \frac{2}{x}$ has a local minimum at
 (A) $x=-2$ (B) $x=0$ (C) $x=1$ (D) $x=2$
84. If $2a+3b+6c=0$, then atleast one root of the equation $ax^2+bx+c=0$ lies in the interval
 (A) (0, 1) (B) (1, 2) (C) (2, 3) (D) (1, 3)
85. At what point on the curve $x^3 - 8a^2y = 0$, the slope of the normal is $-2/3$?
 (A) (2a, -a) (B) (a, a) (C) (2a, a) (D) (-2a, -a)
86. The real number, which most exceeds its cube, is
 (A) 1/2 (B) $\frac{1}{\sqrt{3}}$ (C) $\frac{1}{\sqrt{2}}$ (D) 1/4
87. $\int x e^{x^2} dx$ is equal to
 (A) $-\frac{e^{x^2}}{2} + C$ (B) $\frac{e^{x^2}}{2} + C$ (C) $\frac{e^x}{2} + C$ (D) $-\frac{e^x}{2} + C$
88. If $y = f(x)$ and $y \cos x + x \cos y = \pi$, then the value of $f'(0)$ is
 (A) π (B) $-\pi$ (C) 0 (D) 2π
89. Let $f(\theta) = \sin \theta + \sin 3\theta$ then $f(\theta)$ is
 (A) ≥ 0 only when $\theta \geq 0$ (B) ≤ 0 for all real θ
 (C) ≥ 0 for all real θ (D) ≤ 0 only when $\theta \leq 0$

90. $\int_0^{10\pi} |\sin x| dx$ is
 (A) 20 (B) 8 (C) 10 (D) 18
91. Harish is a voracious reader. The meaning of this sentence is
 (A) Harish hardly reads any books (B) Harish reads lots of books
 (C) Harish reads only one book a day (D) Harish is illiterate
92. The verb in the following sentence is
 "The tiger is a wild animal"
 (A) the (B) tiger (C) is (D) wild
93. The man jumped _____ the fence
 (A) over (B) off (C) out (D) in
94. I accidentally cut _____ with a knife.
 (A) myself (B) herself (C) themselves (D) himself
95. A sudden illness prevented the minister _____ attending the cabinet meeting.
 (A) to (B) from (C) for (D) in
96. I haven't seen him _____ a long time ago.
 (A) since (B) for (C) after (D) from
97. The principal said that the leader of the boys _____ responsible for maintaining the wall magazine.
 (A) can (B) are (C) were (D) was
98. India, of all the countries in the world, _____ the potential to be a frontrunner within the next twenty years.
 (A) has (B) have (C) are having (D) was having
99. I _____ coffee to tea
 (A) like (B) prefer (C) want (D) desire
100. What is the time _____ your watch?
 (A) by (B) in (C) of (D) on